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[Electronegativity - Wikipedia](#)

Electronegativity, symbol χ , is a chemical property that describes the tendency of an atom to attract a shared pair of electrons (or electron density) towards itself.

[ELECTRONEGATIVITY - chemguide](#)

Explains what electronegativity is and how and why it varies around the Periodic Table

[Electronegativity - Chemistry LibreTexts](#)

Electronegativity is a measure of the tendency of an atom to attract a bonding pair of electrons. The Pauling scale is the most commonly used. Fluorine (the most electronegative element) has a Pauling electronegativity of 4.0.

[Electronegativity Definition and Examples - ThoughtCo](#)

Learn the glossary definition of electronegativity, as used in chemistry, chemical engineering, and physics.

[Electronegativity | physics | Britannica.com](#)

Electronegativity: Electronegativity, in chemistry, the ability of an atom to attract to itself an electron pair shared with another atom in a chemical bond. The electronegativity of an atom is influenced by its atomic number and the distance between its nucleus and the electrons shared in the bond.

[Electronegativity - Simple English Wikipedia, the free encyclopedia](#)

Electronegativity, symbol χ , is a chemical property that says how well an atom can attract electrons towards itself. The electronegativity of an atom is influenced by its atomic number and the distance between its nucleus and the electrons shared in the bond.

[Electronegativity and bonding \(video\) | Khan Academy](#)
Electronegativity differences in bonding using the Pauling scale. Classifying bonds as covalent, polar covalent, or ionic.

[Electronegativity - an overview | ScienceDirect Topics](#)

Electronegativity is a chemical property that describes the tendency of an atom or a functional group to attract electrons toward itself. The electronegativity of an

[Learn About Electronegativity and Chemical Bonding](#)
Learn about electronegativity, the trends of electronegativity for elements in the periodic table, and how it relates to chemical bonding.

[Electronegativity and Polar Covalent Bonding - dummies](#)
Electronegativity is the strength an atom has to attract a bonding pair of electrons to itself. When a chlorine atom covalently bonds to another chlorine atom, the electronegativity of the chlorine atom is the same as the electronegativity of the other chlorine atom.

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Electronegativity is the strength an atom has to attract a bonding pair of electrons to itself. When a chlorine atom covalently bonds to another chlorine atom, the electronegativity of the chlorine atom is the same as the electronegativity of the other chlorine atom.

[Electronegativity - polar bonds in organic compounds](#)
An explanation of how electronegativity arises, and the way it produces polar bonds in organic compounds

[Electronegativity | Article about electronegativity by The Editors of Encyclopædia Britannica](#)

electronegativity (Ĭk'trĭnĭg'ĭv'ĭtē), in chemistry, tendency for an atom to attract a pair of electrons that it shares with another atom in a chemical bond.

another atom (see chemical bond

[Electronegativity | definition of electronegativity by ...](#)

Both synchrotron X-ray diffraction and ab initio calculations showed the same cause for bringing the two elements closer in radii and electronegativity, resulting in